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*(See answer key for solutions) If the  $E^\circ$  values of the reaction is negative, then the reaction is NOT spontaneous and therefore the reverse reaction is occurring and the electrons are flowing in the opposite direction. These values are tabulated in the standard reduction potential table. 4. Find*

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$\Delta G$  for the following combined half reactions:  $F_2(g) + 2e^- \leftrightarrow 2F^-(aq)$   
 $Li^+(aq) + e^- \leftrightarrow Li(s) \dots$

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*Key Terms. Make certain that you can define, and use in context, the key term below. hydroboration; Study Notes. The two most*

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*important factors influencing organic reactions are polar (or electronic) effects and steric effects. Hydroboration-Oxidation is a two step pathway used to produce alcohols. The reaction proceeds in an Anti-Markovnikov manner, where the hydrogen (from BH<sub>3</sub> or BHR<sub>2</sub> ...*

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*Answer: According to the correct number of significant figures, we came up with 20.18 amu as the average atomic weight even though the average atomic weight from the periodic table is 20.179 amu.*



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*However, it is still a good check to make sure that you are on the right path. Check Yourself: We predicted earlier that our answer should be closer to the mass of Ne-20 (19.992 amu) instead of Ne-21 ...*

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*Key Takeaways: pH of a Weak Acid. Finding the pH of a weak acid is a bit more complicated than finding pH of a strong acid because the acid does not fully dissociate into its ions. The pH equation is still the same ( $\text{pH} = -\log[\text{H}^+]$ ),*

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*but you need to use the acid dissociation constant ( $K_a$ ) to find  $[H^+]$ . There are two main methods of solving for hydrogen ion concentration. One involves the ...*

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*This example problem demonstrates how to find the energy of a photon from its wavelength. To do this, you need to use the wave equation to relate wavelength to frequency and Planck's equation to find the energy. This type of problem is good practice at rearranging equations, using correct units,*

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*Lewis structures, also known as Lewis dot formulas, Lewis dot structures, electron dot structures, or Lewis electron dot structures (LEDS), are diagrams*

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*that show the bonding between atoms of a molecule and the lone pairs of electrons that may exist in the molecule. A Lewis structure can be drawn for any covalently bonded molecule, as well as coordination compounds.*

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*Similarly, light transmission is a key factor in the ecology of lakes and streams, which are discussed elsewhere ... A View of the Earth, 3rd ed. Englewood Cliffs, NJ: Prentice Hall, 1982. Richardson, Mary Jo, and Wilford D. Gardner. "Tools of the Trade." Quarterdeck 5, no. 1 (1997):10–15. Thurman,*

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Burton. Introductory Oceanography,  
9th ed. Upper Saddle River, NJ ...*

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